# Chapter 2: Matter and Change

**Physical vs. Chemical Change:**

**Physical Change:**

* Properties may change, chemical composition does not (no new substance forms)

Ex. Melt, boil, freeze, condense, break, split, grind, cut, crush

**Chemical Change (reactions):**

* One or more substances combine or breakdown
* Composition (ID) of matter changes, new substance forms

**Reactant ⭢ Product**

**Indicators of Chemical rxn:**

* Formation of a precipitate
* Formation of a gas
* Color change
* Absorption or release of heat
* Formation of Light

**Law of Conservation of Mass:**

* Mass is neither created nor destroyed
* Mass of the reactant(s) = Mass of product(s).

**Practice Problem:**

If 44 grams of carbon dioxide react completely w/ 18 grams of water, what is the mass of carbonic acid formed?

In an engine, octane combines with oxygen to form carbon dioxide and water. If 22.8 grams of octane combine completely w/ 80 grams of oxygen to form 70.4 g carbon dioxide, what mass of water is formed?

**Matter:**

* Anything that has mass and takes up space

**Extensive Properties**:

* Depend on the amt. of matter

Ex. Mass, volume

**Intensive Properties:**

* Depend on the type of matter

Ex. Hardness, Color, solubility, odor, density,

[melting/boiling/freezing points](file:///\\swain.local\storage-ns\homes\high\staff\kgray\Kgray%20folder\Chemistry\chapter%201,2,3-%20intro%20chem,%20matter%20and%20change,%20measurements\melting%20freezing%20point%20example.flipchart)

**Types of Matter:**

**A.) Substance:**

* Uniform, fixed composition can’t separate by physical means
  + - Identical intensive properties

Ex. Gold (Au), Copper (Cu), Water (H20)

**Element:**

* Simplest form of matter -made of only 1 kind of atom
* 100 + known
* Represented by symbols (first letter = capitalized) ex. Na

Ex. Oxygen (O), Hydrogen (H)

**Compound:**

* Substance that contains 2 or more elements **chemically** bonded
* Properties different from their component elements
* Represented by formulas ex. H2O

Ex. NaCl (sodium chloride)

**B.) Mixture:**

* Physical blend of 2 or more components (ex.Salad)
* Composition varies can separate by physical means
* **Phase-** Any part of a sample w/ uniform composition and properties

**Classifying Mixtures:**

**Heterogeneous Mixtures (Suspensions):**

* 2 or more phases (not uniform)

**Homogeneous Mixture (Solutions):**

* 1 phase (uniform)

[**Separating Mixtures**](file:///\\swain.local\storage-ns\homes\high\staff\kgray\Kgray%20folder\Chemistry\chapter%201,2,3-%20intro%20chem,%20matter%20and%20change,%20measurements\Separating%20mixtures.flipchart)**:**

**Filtration:**

* Separating a solid from a liquid in a heterogeneous mixture

**Distillation:**

* Separating a solid from a liquid in a homogenous mixture

**Chromatography**

* Separating homogenous mixtures like ink, pigments

[**States of Matter:**](http://www.youtube.com/watch?v=UnBoQe2rsgo) [**They might be giants**](http://www.youtube.com/watch?v=btGu9FWSPtc)

|  |  |  |  |
| --- | --- | --- | --- |
| **Property** | **Solid** | **Liquid** | **Gas** |
| Shape | Definite | Indefinite | Indefinite |
| **Volume** | Definite | Definite | Indefinite |
| **Expansion** | N/a | Moderate | Great |
| **Compressibility** | N/a | N/a | very |
| **Kinetic Energy** | Low | Medium | High |

13.4 Changes of State

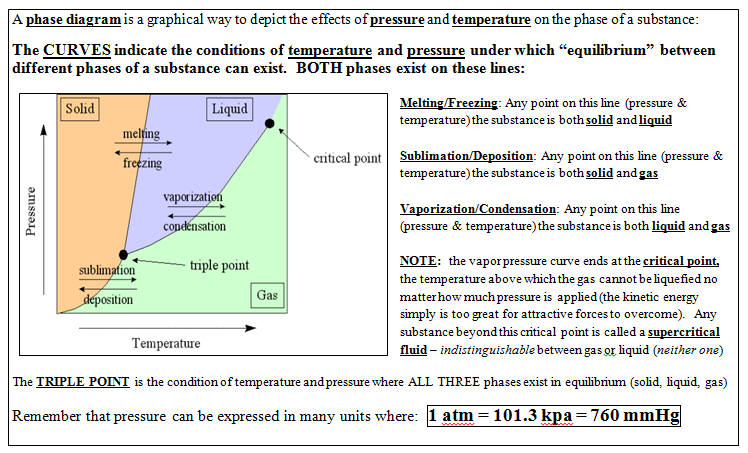
[**Kinetic**](http://www.youtube.com/watch?v=_rsqBNhFG1Y#t=49) **Molecular Theory**

* thermodynamic behavior of matter, especially the relationships among pressure, volume, and temperature in gases
* ⭡ Temperature, ⭡ kinetic energy

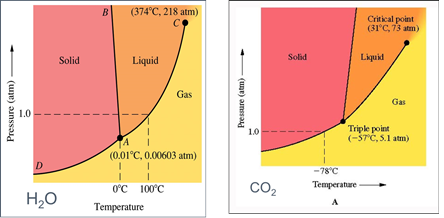
**Entropy:**

* A measure of the disorder of a system.
* Increasing entropy = s → l → g
* [Law](javascript:openGlossaryWnd('e_gchem05_lawdisorder')) of Disorder -states that the natural tendency is for systems to move in the direction of maximum disorder or randomness

**Phase diagram:**



**Water vs. CO2 Phase diagrams**



**\*Sublimation** = solid to gas; atmospheric pressure is too low to stop molecules

from escaping the solid phase; atmospheric pressure less than vapor pressure

\*If the solid phase is **more dense** than the liquid phase. The line that separates

solid and liquids **bends right** (**like in CO2**)

\*IF the solid phase is **less dense** than the liquid phase. The line that separates

solid and liquids **bends left** (**like in water**. So when the pressure increases on

ice it will melt).